Introduction to P-T Equilibrium Phase Diagram

<u>MSE-S203</u> (Phase Equilibria in Materials)

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P-T Equilibrium Phase Diagram

> P-T Phase Diagram

• A substance that has a uniform chemical composition throughout is called a pure substance such as water, air, and nitrogen.

• A pure substance such as water can exist in solid, liquid or vapor phases, depending on the thermodynamic conditions.

• **Example:** Glass of water (liquid phase) with ice cubes (solid phase) in equilibrium.

P-T Equilibrium Phase Diagram

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In a case of Glass of water (liquid phase) containing ice cubes (solid phase) in equilibrium, solid and liquid water are two separate and distinct phases.

Solid and liquid water are separated by a phase boundary (surface of the ice cubes).

• At boiling temperature, liquid water and water vapor are two phases in equilibrium.

PT Equilibrium Phase Diagram for Pure Water



PT Equilibrium Phase Diagram for Pure Water

> P-T Phase Diagram

In the Pressure- Temperature (P-T) phase diagram of pure water, there exist a triple point (solid, liquid and vapor phases of water coexist) at low pressure (0.006 atm) and low temperature (0.0098°C).

• Liquid and vapor phases exist along the vaporization line (phase boundary).

• Liquid and solid phases exist along the freezing line (phase boundary).

P-T Equilibrium Phase Diagram

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• P-T equilibrium phase diagrams also can be constructed for other pure substances.

• **Example:** P-T equilibrium phase diagram for pure iron.

• P-T equilibrium phase diagram for pure iron is different from P-T equilibrium phase diagram for pure water.

• There are three separate and distinct solid phases.

PT Equilibrium Phase Diagram for Pure Iron



PT Equilibrium Phase Diagram for Pure Iron

> P-T Phase Diagram

Alpha Fe, gamma Fe and delta Fe are the three separate and distinct solid phases, present in P-T equilibrium phase diagram for pure iron.

Alpha and delta iron have BCC crystal structure, whereas gamma iron has FCC crystal structure.

Phase boundaries in the solid state have the same properties as the liquid and solid phase boundaries.